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7.B Mole and Mass Avogadro's Number,
The Mole, Grams, Atoms, Molar Mass
Calculations - Introduction Converting
Between Grams and Moles Mole Ratio
Practice Problems Concept of Mole - Part
1 | Atoms and Molecules | Don't
Memorise Empirical Formula \u0026
Molecular Formula Determination From
Percent Composition Molality Practice
Problems - Molarity, Mass Percent, and
Density of Solution Examples ~~Introduction
to Moles~~ Stoichiometry Basic
Introduction, Mole to Mole, Grams to
Grams, Mole Ratio Practice Problems
Mole Conversions Made Easy: How to
Convert Between Grams and Moles How
To Convert Grams To Moles - VERY
EASY! Counting Atoms: Intro to Moles
Part 2 ~~Converting Grams to Moles Using~~

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~~Molar Mass | How to Pass Chemistry The
Mole: Avogadro's Number and
Stoichiometry Introduction to Combustion
Analysis, Empirical Formula \u0026
Molecular Formula Problems Dalton's
Law of Partial Pressure Problems \u0026
Examples - Chemistry Stoichiometry Mole
to Mole Conversions - Molar Ratio
Practice Problems Step by Step
Stoichiometry Practice Problems | How to
Pass Chemistry Balancing Chemical
Equations Practice Problems Concept of
Mole | Avogadro's Number | Atoms and
Molecules | Don't Memorise Chemistry B
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NAME: _____ HR: _____ PAGE 9

CHEMISTRY WORKSHEET # 6

MIXED MOLE PROBLEMS (GRAMS,
MOLECULES, AND LITERS) You now
know three things a mole can be: a molar
mass, 6.02×10^{23} molecules and, for a

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gas, 22.4 liters at STP.

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NAME: HR: _____ PAGE 1

CHEMISTRY B- MOLES PACKET

NAME: _____ HR: _____ PAGE 4

CHEMISTRY WORKSHEET # 2: THE

MOLE AS A UNIT OF MASS Now that

you know how to find the mass of one

mole of a substance (molar mass) you can

easily find the mass of several moles or the

mass of a fraction of a mole using the

factor-label technique.

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A mole of carbon contains 6.02×10^{23}

atoms of carbon. A mole of helium

contains 6.02×10^{23} atoms of helium. A

mole of sodium contains 6.02×10^{23}

atoms of sodium. A mole of gold contains

6.02×10^{23} atoms of gold. However, we

have learned about seven elements that

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exist as diatomic molecules—H₂, N₂, O₂, F₂, Cl₂, Br₂, and I₂

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THE MOLE AS A UNIT OF MASSNow

that you know how to find the mass of one

mole of a substance (molar mass) you can

easily find the mass ofseveral moles or the

mass of a fraction of a mole using the

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factor-label technique. 1 mol = a molar mass of an atom/molecule (g/mol) EXAMPLE: What is the mass of 5.00 moles of water H₂O? STEP 1: 2 H = 2 x (1.01) = 2.02 STEP 2: ? grams H₂O = 5.00 moles H₂O x 18.02 g H₂O = 90.10 ...

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NAME: _____ HR: _____ PAGE

6 CHEMISTRY WORKSHEET # 4

MOLAR MASSES AND

AVOGADRO ' S NUMBER Now that

you know two definitions of a mole (a molar mass and an Avogadro ' s number of particles) you can combine these two definitions into one problem. EXAMPLE:

How many molecules are there in 90.1 grams of water? 2 H = 2 x (1.01) = 2.02 O = 1 x (16.00) = 16.00 H₂O = 18.02 g? molecules H₂O = 90.1 g H₂O x 1 mole

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$H_2O \times 6.02 \times 10^{23} \text{ molecules } H_2O =$
 $3.01 \times 10^{24} \text{ molecules } H_2 \dots$

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WORKSHEET # 6 MIXED MOLE
PROBLEMS (GRAMS, MOLECULES,

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AND LITERS) You now know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4 liters at STP. We can use this

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Question: PAGES CHEMISTRY B-
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CHEMISTRY WORKSHEET #5
MOLE PROBLEMS-MOLAR
VOLUME OF A GAS We Have Learned

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That A Mole Is A Mass Of Material And Number Of Particles. A Mole Can Also Be A Measure Of Volume When We Are Talking About Gases. You May Remember From Previous Science Classes That All Gases Behave Basically The Same As Far As The Physical ...

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solutions for you to be successful. Page 6/9

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Page 12/13

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Stoichiometry (7(www.mrpalermo.com(Use the formula below to answer questions 8-13!
$$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$$
 (8.
If 3.00 moles of Iron ...

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